

Using the last result we can calculate

$$a = a_x = -a_y = \frac{m_1}{m_1 + m_2} g,$$

$$T_2 = T_1 = \frac{m_2 m_1}{m_1 + m_2} g.$$

Numerical results:

$$a = a_x = \frac{3}{5} \cdot 9.81 \text{ m s}^{-2} = 5.89 \text{ m s}^{-2},$$

$$T_1 = T_2 = 1.18 \text{ N}.$$

**Problem 2.** Water of mass  $m_2$  is contained in a copper calorimeter of mass  $m_1$ . Their common temperature is  $t_2$ . A piece of ice of mass  $m_3$  and temperature  $t_3 < 0^\circ\text{C}$  is dropped into the calorimeter.

- a) Determine the temperature and masses of water and ice in the equilibrium state for general values of  $m_1$ ,  $m_2$ ,  $m_3$ ,  $t_2$  and  $t_3$ . Write equilibrium equations for all possible processes which have to be considered.
- b) Find the final temperature and final masses of water and ice for  $m_1 = 1.00 \text{ kg}$ ,  $m_2 = 1.00 \text{ kg}$ ,  $m_3 = 2.00 \text{ kg}$ ,  $t_2 = 10^\circ\text{C}$ ,  $t_3 = -20^\circ\text{C}$ .

Neglect the energy losses, assume the normal barometric pressure. Specific heat of copper is  $c_1 = 0.1 \text{ kcal/kg}\cdot^\circ\text{C}$ , specific heat of water  $c_2 = 1 \text{ kcal/kg}\cdot^\circ\text{C}$ , specific heat of ice  $c_3 = 0.492 \text{ kcal/kg}\cdot^\circ\text{C}$ , latent heat of fusion of ice  $l = 78,7 \text{ kcal/kg}$ . Take  $1 \text{ cal} = 4.2 \text{ J}$ .

*Solution:*

We use the following notation:

$t$	temperature of the final equilibrium state,
$t_0 = 0^\circ\text{C}$	the melting point of ice under normal pressure conditions,
$M_2$	final mass of water,
$M_3$	final mass of ice,
$m'_2 \leq m_2$	mass of water, which freezes to ice,
$m'_3 \leq m_3$	mass of ice, which melts to water.

- a) Generally, four possible processes and corresponding equilibrium states can occur:

1.  $t_0 < t < t_2$ ,  $m'_2 = 0$ ,  $m'_3 = m_3$ ,  $M_2 = m_2 + m_3$ ,  $M_3 = 0$ .

Unknown final temperature  $t$  can be determined from the equation

$$(m_1c_1 + m_2c_2)(t_2 - t) = m_3c_3(t_0 - t_3) + m_3l + m_3c_2(t - t_0). \quad (7)$$

However, only the solution satisfying the condition  $t_0 < t < t_2$  does make physical sense.

2.  $t_3 < t < t_0$ ,  $m'_2 = m_2$ ,  $m'_3 = 0$ ,  $M_2 = 0$ ,  $M_3 = m_2 + m_3$ .

Unknown final temperature  $t$  can be determined from the equation

$$m_1c_1(t_2 - t) + m_2c_2(t_2 - t_0) + m_2l + m_2c_3(t_0 - t) = m_3c_3(t - t_3). \quad (8)$$

However, only the solution satisfying the condition  $t_3 < t < t_0$  does make physical sense.

3.  $t = t_0$ ,  $m'_2 = 0$ ,  $0 \leq m'_3 \leq m_3$ ,  $M_2 = m_2 + m'_3$ ,  $M_3 = m_3 - m'_3$ .

Unknown mass  $m'_3$  can be calculated from the equation

$$(m_1c_1 + m_2c_2)(t_2 - t_0) = m_3c_3(t - t_3) + m'_3l. \quad (9)$$

However, only the solution satisfying the condition  $0 \leq m'_3 \leq m_3$  does make physical sense.

4.  $t = t_0$ ,  $0 \leq m'_2 \leq m_2$ ,  $m'_3 = 0$ ,  $M_2 = m_2 - m'_2$ ,  $M_3 = m_3 + m'_2$ .

Unknown mass  $m'_2$  can be calculated from the equation

$$(m_1c_1 + m_2c_2)(t_2 - t_0) + m'_2l = m_3c_3(t_0 - t_3). \quad (10)$$

However, only the solution satisfying the condition  $0 \leq m'_2 \leq m_2$  does make physical sense.

b) Substituting the particular values of  $m_1$ ,  $m_2$ ,  $m_3$ ,  $t_2$  and  $t_3$  to equations (7), (8) and (9) one obtains solutions not making the physical sense (not satisfying the above conditions for  $t$ , respectively  $m'_3$ ). The real physical process under given conditions is given by the equation (10) which yields

$$m'_2 = \frac{m_3c_3(t_0 - t_3) - (m_1c_1 + m_2c_2)(t_2 - t_0)}{l}.$$

Substituting given numerical values one gets  $m'_2 = 0.11$  kg. Hence,  $t = 0^\circ\text{C}$ ,  $M_2 = m_2 - m'_2 = 0.89$  kg,  $M_3 = m_3 + m'_2 = 2.11$  kg.